Name: _____	**Subject:** Chemistry	**Class:** 10th	**Time:** 60 minutes	Total Marks:	**30**
Chapter No.	**MJDexpert.com**			Obtained marks	

Note: Please attempt any 7 short questions from Question 2. Also, attempt both parts (a and b) of Question 3. "a" Part is 5 marks and "b" is 4 marks. Cutting and removal of any content is strictly prohibited.

Objective-Section

Q. 1 Encircle the correct Answer . (7x1=7)

1. The colour of HI in reversible reaction is

A) Blue B) purple C) Green D) Colourless

2. When. Rate of forword reaction =Rate of reverse reaction. This condition is called

A) Dynamic Equilibrium state B) static equilibrium state C) constant equilibrium state D) All of these

3. Forward reaction gradually

A) fast B) slow C) moderate D) hight

4. If Qc <Kc then the direction of reaction is

A) constant B) forword C) reverse D) none of these

5. At Dynamic Equilibrium

A) The reaction stops to proceed B) The amount of reactant and product are equal

C) Establishes after completion of reaction D) establishes readily

6. The Kc value depond upon

A) amount of Reactant B) Amount of products C) Catalyst D) Temprature

7. A reverse reaction is one that

a) which proceeds from left to right b) in which reactants react to form products

c) which slow down gradually d) which speeds up gradually

Subjective-Section

Q.2 Write short answers of any seven of the following questions: (7x2=14)

i. Write down the difference between reversible and irreversible reaction?
ii. Write down the difference between forword reaction and reverse reaction?
iii. Define static equilibrium with example ?
iv. What happened when Qc > Kc ?
v. What happens when Kc value is large?
vi. Write down the characteristics of reversible reaction?
vii. Explain reversible reaction with Graph ?
viii. Define equilibrium constant?

Q.No.3 Long Question: (5+4=9)

a) Write down the macroscopic properties of dynamic equilibrium?
b) Explain law of mass action?