

Name: _____						Subject: Chemistry		Class: 9 <sup>th</sup>		Time: 60 minutes		Total Marks: 30	
Chapter No.4				MJDexpert.com						Obtained marks			

**Note:** Please attempt any 7 short questions from Question 2. Also, attempt both parts (a and b) of Question 3. "a" Part is 5 marks and "b" is 4 marks. Cutting and removal of any content is strictly prohibited.

## Objective-Section

**Q. 1 Encircle the correct Answer .**

**(7x1=7)**

### MJD Chemistry

1.	The electrons in the outer most shell of an atom are called:														
a	Valence electron			b	Inner electron			c	Lone pair			d	None of these		
2.	The bond formation between ions is due to:														
a	Electron sharing			b	Inter molecular forces			c	Electro-static forces			d	Repulsive forces		
3.	The number of electrons in valance shell of chlorine is _____.														
a	6			b	7			c	5			d	4		
4.	Atomic number of sodium is:														
a	11			b	10			c	12			d	13		
5.	The bond formed due to mutual sharing of electron is _____.														
a	Metallic bond			b	Ionic bond			c	Co-ordinate bond			d	Covalent bond		
6.	The formation of ammonium ion is due to:														
a	Covalent bond			b	Ionic bond			c	Metallic bond			d	Co-ordinate Covalent bond		
7.	The atom which donate electron pair in dative covalent bond is know:														
a	Acceptor			b	Donor			c	Electronegative			d	Ionic bond		

## Subjective-Section

**Q.2 Write short answers of any seven of the following questions: (7x2=14)**

1. Define metallic bond?
2. Define intermolecular forces? Show these forces among HCl molecule?
4. What type of bond exists in sodium chloride?
5. Metals are good conductors of electricity why?
6. Why do atoms reacts?
7. Ionic compounds are solids justify?
8. Differentiate between lone pair and bond pair of electrons?

**Q.No.3 Long Question:**

**(5+4=9)**

- a) Define ionic bond and explain the formation of NaCl.
- b) Define hydrogen bonding with example and explain the effect of hydrogen bonding on physical properties.