

Name: _____						Subject: Chemistry		Class: 9 <sup>th</sup>		Time: 60 minutes		Total Marks: 30	
Chapter No.3				MJDexpert.com				Obtained marks					

**Note:** Please attempt any 7 short questions from Question 2. Also, attempt both parts (a and b) of Question 3. "a" Part is 5 marks and "b" is 4 marks. Cutting and removal of any content is strictly prohibited.

### Objective-Section

**Q. 1 Encircle the correct Answer . (7x1=7)**

1.	The horizontal rows of elements in periodic are called?											
a	Groups	b	Periods	c	Blocks	d	Shell					
2.	The vertical columns of elements in periodic table are called:											
a	Groups	b	Periods	c	Blocks	d	Rows					
3.	4 <sup>th</sup> and 5 <sup>th</sup> period of long form of period table are called:											
a	Short periods	b	Normal periods	c	Long periods	d	Very long periods					
4.	2nd 3rd periods of long form of periodic table are called _____?											
a	Short periods	b	Normal periods	c	Long periods	d	Very long periods					
5.	6 <sup>th</sup> and 7 <sup>th</sup> periods of long form of periodic table are called?											
a	Short periods	b	Normal periods	c	Long periods	d	Very long periods					
6.	Alkaline earth metals belong to _____ group.											
a	First	b	Second	c	Third	d	Fourth					
7.	The distance between nuclei of two carbon atoms is:											
a	152 pm	b	153 pm	c	154 pm	d	155pm					

### Subjective-Section

**Q.2 Write short answers of any seven of the following questions: (7x2=14)**

- Why noble gases are not reactive?
- Write down the difference between Dobereiner triads & Newland octaves?
- Write down the difference between Mendeleev's periodic law and Modern periodic law?
- Write down the difference between group and period?
- Define Atomic size with example ?
- Define shielding effect?
- Define Electronegativity with example ?
- Why the size of an atom does not decrease regularly in a period?

**Q.No.3 Long Question: (5+4=9)**

- Write a detail note on atomic radius?
- Define electron affinity why it increases in a period and decrease in group in periodic table?