

Name: _____						Subject: Chemistry		Class: 9 th		Time: 60 minutes		Total Marks: 30	
Chapter No.2				MJDexpert.com				Obtained marks					

Note: Please attempt any 7 short questions from Question 2. Also, attempt both parts (a and b) of Question 3. "a" Part is 5 marks and "b" is 4 marks. Cutting and removal of any content is strictly prohibited.

Objective-Section

Q. 1 Encircle the correct Answer .

(7x1=7)

MJD Chemistry

1.	The scientist who put plum pudding theory is _____?														
a	Bohar			b	Thomson			c	Chadwick			d	Rutherford		
2.	The mass of an electron is:														
a	1.672×10^{-24} kg			b	1.672×10^{-24} g			c	9.106×10^{-28} kg			d	9.106×10^{-28} g		
3.	In discharge tube the canal rays are produced due to _____?														
a	Presence of anode			b	Ionization of gas molecules			c	Presence of cathode			d	High pressure of gas		
4.	The value of plank's constant is?														
a	6.63×10^{-36} js			b	6.63×10^{-33} js			c	6.63×10^{-34} js			d	6.63×10^{-33} js		
5.	Who is considered as father of Nuclear science?														
a	Rutherford			b	Planck			c	Bohr			d	Thomson		
6.	Bohr won the noble prize in?														
a	1922			b	1923			c	1920			d	1921		
7.	The electronic configuration if hydrogen is:														
a	$1s^2, 2s^2$			b	$1s^2$			c	$1s^2, 2s^1$			d	$1s^1$		

Subjective-Section

Q.2 Write short answers of any seven of the following questions: (7x2=14)

- i. Give four characteristics of cathode rays?
- ii. Write down the properties of canal rays?
- iii. Define nucleons ?
- iv. Write down the defects of Rutherford Atomic model?
- v. Write down the difference between Rutherford Atomic theory and Bohr Atomic theory?
- vi. Define Electronic configuration with example ?
- vii. Write down the electronic configuration of Cl & Na?
- viii. Write down the two uses of isotopes?

Q.No.3 Long Question:

(5+4=9)

- a) Explain Bohr Atomic theory?
- b) Define isotopes with examples?